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Assessing the Credibility of the Calorific Value of Municipal Solid Waste

U.S. DEPARTMENT OF COMMERCE
National Bureau of Standards
Chemical Thermodynamics Division
Center for Chemical Physics

and the

National Measurement Laboratory*

and the

Temperature and Pressure Division**
Center for Basic Standards
Washington, DC 20234

February 1984

Prepared for

U.S. Department of Energy
Office of Renewable Technology
Energy from Municipal Waste Division
Forrestal Building

1000 Independence Avenue, SW
Washington, DC 20585

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K. L. Churney, E. S. Domalski, A. E. Ledford, J. C. Colbert,
S. S. Bruce, T. J. Buckley, R. C. Paule*, and M. L. Reilly**

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U.S. DEPARTMENT OF COMMERCE, Malcolm Baldrige, *Secretary*
NATIONAL BUREAU OF STANDARDS, Ernest Ambler, *Director*

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PREFACE

This report gives the results of studies which were started in fiscal year 1981, continued through fiscal year 1982, and were completed in fiscal year 1983. Two important industrial/municipal queries required response in this study. In response to the first, we compared the precision and accuracy of the calorific value of municipal solid waste (MSW) which had been processed to 2 mm particle size or less and which had been burned in both a gram-size bomb calorimeter and in the kilogram-size flow calorimeter. In response to the second query, we determined whether there would be a significant difference between the results of burning an unprocessed and processed MSW sample in the kilogram-capacity combustion flow calorimeter. The results of the calorimetric studies which respond to these queries are presented in this report.

The chronology of progress in this study was as follows:

Oct 80 - Feb 81	Design of kilogram capacity flow calorimeter.
Feb 81 - Apr 81	Critique and development of engineering drawings.
Apr 81 - Jun 81	Competitive bidding.
Jun 81 - Jul 81	Awarding of bid to vendors.
Jul 81 - Feb 82	Fabrication period.
Feb 82 - Apr 82	Assembly.
Apr 82 - Dec 82	Preliminary testing, combustion runs (First combustion May 82).
Jan 1983	Baltimore County Sampling.

Feb 83	Development of Measurement Protocol (Statistically designed experiments).
Feb 82 - May 83	Construction and completion of product gas analysis train.
Apr 83 - Oct 83	Calorimetric Measurements. (A total of 26 calorimetric experiments were carried out).
Sep 13, 1983	DoE Research Review on the Characterization of Municipal Solid Waste.
Sep 83 - Oct 83	Calculation of Results.

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Abstract

A study has been made at the National Bureau of Standards to establish the limits of reliability of the calorific value of municipal solid waste (MSW) determined by the bomb calorimetric procedure currently used in commercial test laboratories. This procedure involves using gram-size samples derived from MSW that has been processed down to a particle size of 2 mm or less. Critics of the procedure argue that gram-size samples are too small to be representative of such a large quantity of so heterogeneous a material, and that processing MSW may also alter its composition.

To test the bomb calorimetric procedure, a 2.5 kg capacity combustion flow calorimeter was designed and constructed for the determination of the enthalpies of combustion of kilogram-size samples of MSW in flowing oxygen near atmospheric pressure.

Calorimetric data on processed MSW were obtained using both the kilogram-size flow and a gram-size bomb calorimeter. Intercomparison of results shows that the calorific value (on a dry basis) of gram-size test samples agrees, within the uncertainty of our experiments, with the corresponding values for their kilogram-size parent samples provided that the sample division technique used to obtain the gram-size samples is that described in this work. The average difference of the parent minus gram-size sample values (on a dry basis)

is -0.1% with an imprecision (95% confidence interval) of $\pm 1.1\%$ of the mean calorific value.

The effects of processing on sample composition were determined by inter-comparison of flow calorimetric results on kilogram-size samples of processed and minimally processed MSW (150 mm or less particle size) that are nominally identical. The average difference of the unprocessed minus processed values (on a dry basis) is -0.5% with an imprecision (95% confidence interval) of $\pm 2.9\%$ of the mean calorific value.

1. Introduction

The work described in this paper was performed to assist in resolving an industrial problem formulated by members of the ASME Research Committee on Industrial and Municipal Wastes, the ASME Performance Test Code Committee (PTC-33) on Large Incinerators, and the ASTM Committee (E-38) on Resource Recovery. Incinerator-boiler systems which are used to dispose of municipal solid waste (MSW) are bought and sold on the basis of thermal specifications. To determine whether a system meets its thermal specifications, the calorific value of the input waste stream fed to the system must be known. At present, the calorific value is determined at commercial test laboratories using gram-size samples of the waste.

It has been the opinion of many combustion engineers that one cannot accurately sample a large, multiton quantity of MSW down to one or two grams and obtain a representative test sample. The calorific value determined from such a small sample would be suspect. Gram-size samples must be prepared from the bulk MSW by separation and milling techniques to provide a powdered test sample of approximately 2 mm or less particle size. Many combustion engineers feel that this processing may significantly alter the composition of the test sample and, hence, the calorific value. This might occur because of changes in composition

due to excessive local heating associated with some milling operations or due to sampling and handling a material which is inhomogeneous and segregates easily.

The general question of whether or not the calorific value of a multiton quantity of MSW can be obtained by sampling down to one or two grams can be broken into two parts. The first part is: How should one sample a multiton pile or stream of raw waste in order to obtain increments of sufficient number and size to characterize the calorific value with a given precision? The second part is: How can one obtain representative gram-size samples that characterize these increments? The various issues in the dispute about the validity of the current method for determining the calorific value of MSW are primarily concerned with the second of these questions.

To address these issues, the National Bureau of Standards (NBS) initiated a research program in collaboration with the U.S. Department of Energy to develop test procedures for accurately determining the calorific values of refuse and refuse-derived fuels (RDF). This research also was part of the NBS response to the Resource Conservation and Recovery Act of 1976. Subtitle E of this legislation mandated NBS to provide guidelines for the development of specifications for classification of materials destined for disposal.

NBS made a commitment to develop a calorimeter to accommodate a sample of MSW which is larger than that currently used in commercial test laboratories. In addition the MSW sample should be processed as little as possible. After many discussions with combustion engineers and statisticians, we decided that an acceptable sample size was about 2.5 kg and that the sample would be derived from dried RDF-2⁽⁺⁾. In terms of the general sampling question, cited

⁽⁺⁾ RDF-2 is MSW that has been processed to reduce the particle size so that 95 mass-percent passes through a 15 cm square mesh screen.

previously, the sample size adopted for the calorimeter is assumed to be a convenient fraction of a sampling increment for the waste stream.

For safety considerations, we chose to build a calorimeter for the combustion of samples in flowing oxygen near atmospheric pressure rather than to scale up the conventional high pressure oxygen combustion bomb. After the 1880's, the development of the flow technique was discontinued in favor of the simpler procedures and more quantitative results which could be obtained with the bomb calorimeter. As a consequence, the first goal of this project was to demonstrate that the oxygen flow technique could be used to obtain complete combustion of MSW.

To test the oxygen flow technique, a combustion flow calorimeter [1,2] was built and used to successfully combust 25 g pellets of RDF-4⁽⁺⁾. Combustion of RDF which has been compressed into a pellet rather than left in loose form was adopted because this configuration simplifies quantitative collection of ash and prevents dispersion of the sample away from the main reaction zone. The amount of carbon in the ash and the uncertainty in the amount of CO in the combustion products contributed errors of less than 0.1% to the enthalpy of combustion.

After an initial study of the variability of MSW was carried out at NBS in 1980 [3], it was concluded that a calorimeter which could burn a kilogram-size sample with a total uncertainty of less than one percent in the enthalpy of combustion would satisfy our requirements.

Prior to designing a large-scale combustion flow calorimeter, the burning characteristics of kilogram-size pellets of RDF-2 in oxygen were determined [4].

(+) RDF-4 is MSW from which the metal has been removed and then processed to reduce the particle size so that 95 mass-percent passes through a 2 mm square mesh screen.

The burning characteristics were unknown and were expected to be different from RDF-4. The physical heterogeneity, measured against sample size, and variability of composition of RDF-2 pellets are greater than those of RDF-4 pellets. The goal of these experiments was to develop a method of controlled burning of a 2 to 2.5 kg pellet at a rate of 15 minutes or less per kilogram. This time limit was the estimated maximum time which could guarantee that the imprecision contributed by the calorimetric measurement would be less than one percent.

A series of eighteen trial experiments were carried out in a prototype combustor [4]. Burning times of 15 min/kg for RDF-2 were obtained using the following measures. The sample was enclosed by a crucible to reduce heat loss. A small, secondary, flow of oxygen in the form of horizontal jets of preheated oxygen was directed at the sides of the cylindrical pellet. The larger, primary, oxygen flow was supplied as a diffuse flow that was directed at the bottom of the sample.

After developing a suitable technique for burning a kilogram-size sample of RDF-2, a 2.5 kg capacity flow calorimeter was designed, fabricated, and placed in operation. Calorimetric measurements have been carried out in a study consisting of 26 experiments designed to resolve part of the dispute about the current method of determining the calorific value of MSW. A brief description of the calorimeter, auxiliary measurement equipment, and the main features of a typical experiment is given in section 2. The measurement protocol is discussed in section 3. Results are analyzed in section 4. Conclusions are given in section 5.

2. Experimental Apparatus

2.1 Description of 2.5 Kg Flow Calorimeter and Auxiliary Apparatus

The flow calorimeter consists of two basic parts: a constant temperature jacket and the calorimeter proper. The latter contains the sample combustor.

Heat liberated in the combustion reaction is transferred to the stirred water in the calorimeter vessel, the outermost container of the calorimeter proper. The rise in temperature of the calorimeter water, after correction for heat exchange with the jacket and for heat transport by the flowing gases, is proportional to the enthalpy of combustion of the sample. The calorimeter is calibrated by determining the temperature rise produced by combustion of a known mass of microcrystalline cellulose whose enthalpy of combustion has been established by bomb calorimetry.

A cross section of the flow calorimeter is shown in Figure 1. The important dimensions and calorimeter properties are listed in Table 1. All metal components were fabricated from type 316 stainless steel.

The cylindrical compressed sample pellet (nominally 23 cm diameter and 15 cm high), A, is located inside and near the bottom of the combustor, B. The pellet is supported on a horizontal lattice of nine alumina rods (6 mm diameter) resting in notches cut in the top edge of a cup-shaped ash pan. The ash pan sits on the access plate that forms the bottom of the combustor and retains the residual ash from the burned MSW. An iron fuse wire, having a four turn coil touching the top of the sample at the center, is stretched horizontally between two electrodes (shown on either side of the sample in Figure 1). The sample is ignited by passage of electrical current through the fuse wire.

Flowing oxygen is supplied to the sample by an array of nozzles arranged in five horizontal tiers. Each tier contains six symmetrically spaced nozzles located in the combustor wall. The bottom tier supplies oxygen through slots in the ash pan to form a diffuse stream that is directed at the bottom of the sample. The other tiers supply narrow jets of oxygen that are either directed at the sides of the sample or into the space above the sample. The oxygen of the top four tiers is preheated by passing it through coils wound around and

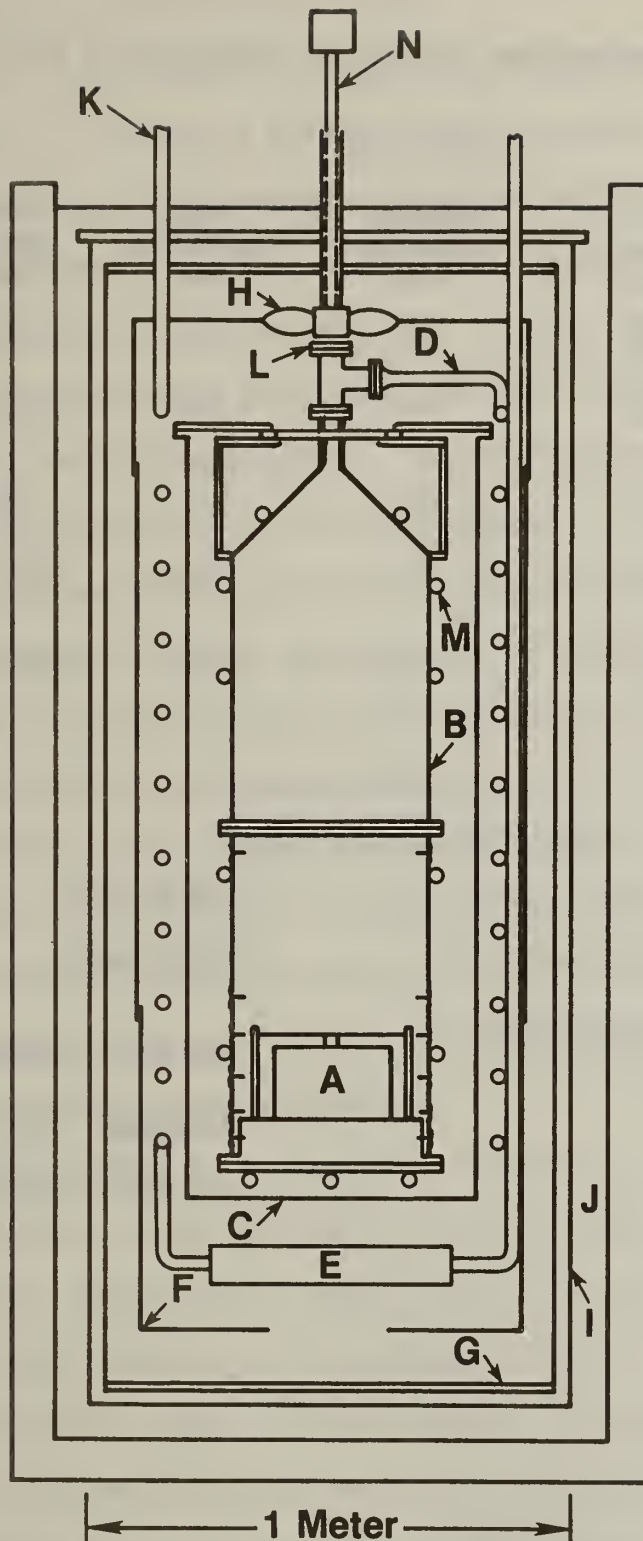


Figure 1. Diagram of the 2.5 kg capacity flow calorimeter. A denotes the sample pellet, B the combustor, C the combustor enclosure, D the exhaust coil, E the collector, F the flow shield, G the calorimeter vessel, H the stirrer, I the submarine jacket, J the jacket water, K a quartz thermometer, L the window, M the quick-cool heat exchange system, and N the boroscope.

Table 1. Specifications for the 2.5 kg capacity flow calorimeter.

Component	Dimensions			
	Height (cm)	Outside Diameter (cm)	Wall Thickness (mm)	Total Mass (kg)
Combustor	155	41	3.2	74
Combustor Enclosure	163	61	4.8	196
Flow Shield	213	79	1.6	73
Calorimeter Vessel	236	91	4.8	390
Submarine Vessel	243	97	4.8	454

Additional Specifications

Calorimeter water	980 liters
Jacket water	2160 liters
Heat capacity of calorimeter	0.4 MJ/K (assembly)
	<u>4.1 MJ/K</u> (water)
	4.5 MJ/K (total)

welded to the outside of the combustor wall.

The combustor is surrounded by a concentric cylindrical vessel, the combustor enclosure, C. The space between B and C contains tubes that supply oxygen to the combustor, thermocouples mounted on the combustor wall, tubes (M) of a quick-cool heat exchange system (also welded to the outside of the combustor), and argon gas to reduce corrosion of these components.

The product gas leaves the top of the combustor and then passes downward through a ten-turn heat-exchange coil, D, then through the collector, E, and finally through the exit tube out of the calorimeter. The heat exchange coil transfers heat from the product gas to the calorimeter water. Most of the water formed in the combustion reaction condenses and is trapped in the collector. The calorimeter water is directed past the heat exchange coils by means of a flow shield, F. Water is drawn through the bottom port (20 cm diameter) in the flow shield, past the coils, and is mixed by stirrer blades, H, mounted in the top port of the shield. Water returns to the bottom port via the annular space between the calorimeter vessel, G, and flow shield at a nominal water flow rate of 380 liters per minute (lpm). Thus, the calorimeter water makes a complete circuit of the flow path every 2.5 minutes.

The calorimeter vessel is surrounded by a concentric cylindrical vessel, the submarine vessel I, which constitutes the innermost portion of the constant temperature jacket. The 2.5 cm air space between the vessels reduces the heat exchange between the jacket and the calorimeter proper. The submarine vessel is submerged in the stirred water, J, of the jacket. The calorimeter water temperature is monitored by a quartz oscillator thermometer, K. A similar thermometer monitors the temperature of the jacket water. Temperatures of the inlet oxygen supply and the product gas are monitored with thermocouple probes whose junctions are located in the gas tubes at the level of the calorimeter

vessel lid.

The combustion reaction is monitored visually through the window, L, sealed to the inline port of a tee in the product gas line (at the top of the combustor in Figure 1) by means of the boroscope, N, which extends through the hollow stirrer shaft. The boroscope is attached to a TV camera which is connected to a video monitor-recording system.

From the calorimeter vessel, the product gas passes through a mixing chamber and flowmeter. Just beyond the flowmeter, a portion of the gas is continuously withdrawn for analysis. Part of this gas is dried and passed through dedicated infrared detectors for CO and CO₂. A second part is passed through a scanning infrared detector used to monitor various trace components (hydrocarbons, SO₂, HCl, etc.). The remaining part of this gas is passed through a cooled-mirror type of automatic dew point detector to monitor water vapor. These detectors in combination with three manometers and six thermocouples (to monitor gas temperatures) are used to determine the composition of the product gas.

A dual minicomputer based data logging and computation system is used to monitor the outputs of a total of 33 thermocouples, two quartz thermometers, eight flowmeters, and seven instruments in the product gas analysis train.

2.2 Description of a Typical Experiment

Before ignition of the sample, the calorimeter water temperature is monitored for one hour after the drift rate becomes constant. During this initial drift period and in a similar length final drift period after the combustion reaction, the calorimeter water is pumped through the quick-cool heat exchange tubes (M of Figure 1) in the combustor wall by a gear pump mounted in the calorimeter water. This increases the rate of temperature equilibration of the combustor with the calorimeter water and insures that the calorimeter proper

is more nearly isothermal during the drift periods.

Ten minutes prior to ignition, the supply of water to the quick-cool heat exchange system is turned off (i.e. water from the gear pump is returned directly to the calorimeter vessel) and the water remaining in the quick-cool system is blown out with air. The combustor is flushed with oxygen. During the combustion, oxygen is supplied from a manifold containing eight 6200 liter (STP) standard oxygen tanks and six mass flow controllers. Oxygen is supplied to the combustor, the underside of the viewing window (to prevent water condensation), and the product gas as it leaves the collector. The last supply reduces the dew point of the product gas below room temperature in order to prevent water from condensing in the analysis train. Typical flow rates are 150 lpm to the combustor, 10 lpm to the viewing window, and 110 lpm to dilute the product gas for a total flow of 270 lpm.

Sample burning times were approximately 32 minutes for RDF-2 and micro-crystalline cellulose, and about 42 minutes for RDF-4. The CO_2 , H_2O , and CO concentrations in the product gas during the combustion of an RDF-4 pellet are shown in Figure 2. The peak and average CO_2 concentrations over the first ten minutes of the combustion are 52 and 38 mole percent, respectively. The latter concentration corresponds to an oxygen supply rate of about 2.5 times stoichiometry. The H_2O concentration reaches a maximum of 2.4 mole percent (i.e. dew point of 16 °C) at 30 minutes; the abrupt fall in concentration at 42 minutes indicates that the combustor has been completely flushed of product gas. The occurrence of two peaks in the CO concentration versus time is typical of all the RDF samples. In Figure 2, the initial and final maximum CO concentrations are 0.016 and 0.027 mole percent, respectively. The initial peak always occurs near the time of the maximum rate of production of CO_2 ; the final peak occurs near the end of the visual burning when the ash tends to inhibit the combustion.

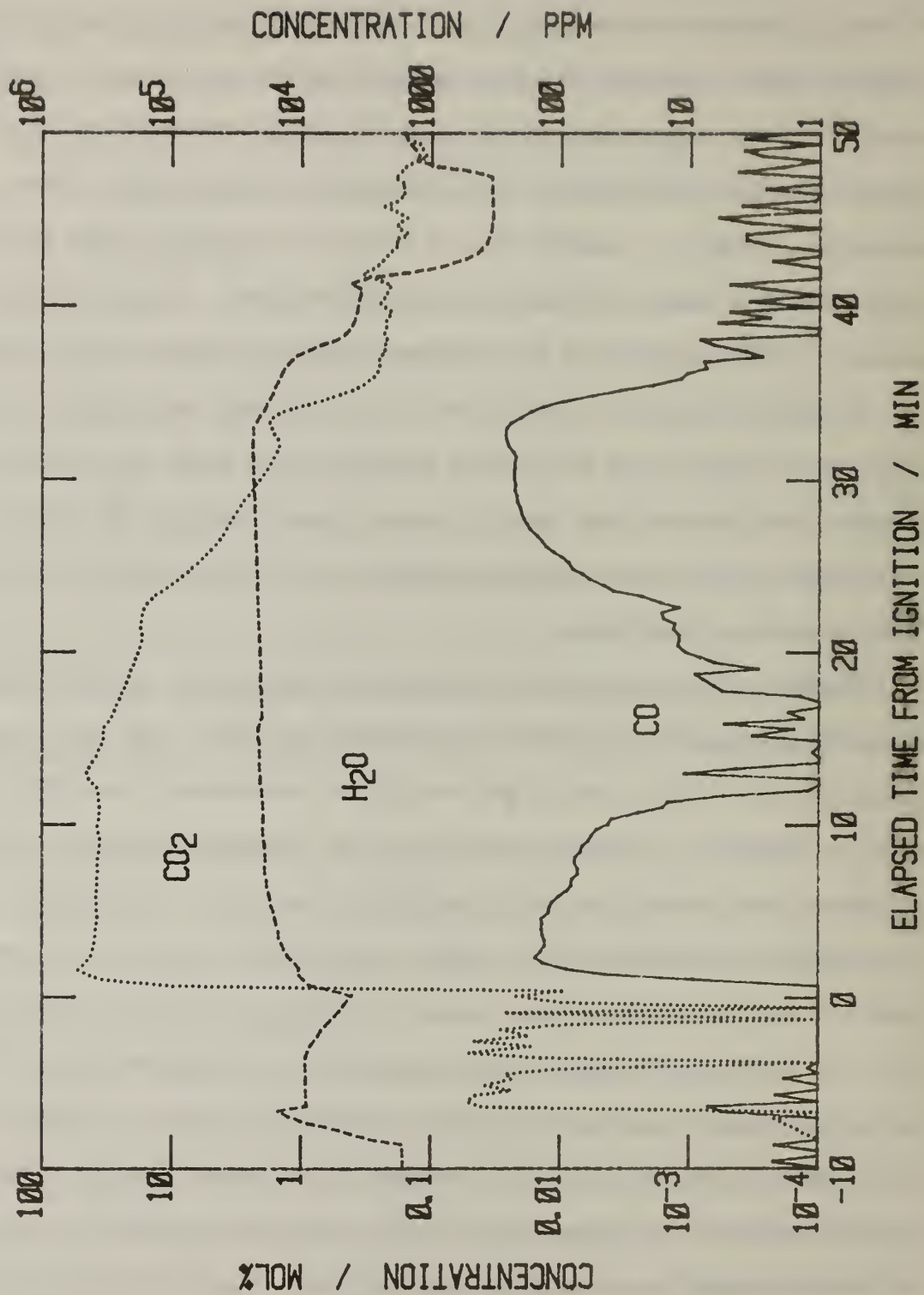


Figure 2. A plot of the concentrations of gaseous carbon dioxide, water, and carbon monoxide as a function of time for a typical experiment on RDF.

The average CO concentration in a cellulose combustion is about a factor of fifty smaller than that shown in Figure 2.

The temperature of the product gas as it leaves the top of the combustor rises to about 400 °C in the first minute of the combustion, gradually increases to a maximum of about 450 °C between 12 and 15 minutes after ignition, and then decays exponentially thereafter. The temperatures of the combustor wall and of the oxygen gas as it leaves the preheat coils increase more gradually to peak values at about the same time and decay similarly. The peak values of the temperatures of the wall and adjacent preheat coils within 25 cm of the ash pan are between 600 and 700 °C.

When the temperature of the combustor base plate has decreased to 220 °C, the calorimeter water is again circulated through the quick-cool heat exchange system. This safely removes the appreciable excess heat stored in the combustor. A plot of the calorimeter water temperature vs elapsed time for the experiment corresponding to Figure 2 is shown by the solid line of Figure 3. The temperature of the calorimeter water had increased 8 °C after an elapsed time of 60 minutes when the quick-cool system was activated. During the next 10 minutes, when most of the excess heat was removed, the temperature of the calorimeter water increased by 1.2 °C. This increase corresponded to 13% of the total heat liberated by the combustion reaction.

After the drift rate of the calorimeter water attains a steady value, temperature readings are made for an additional hour which corresponds to the final drift period. The data points in Figure 3 are the deviations of the calorimeter water temperature from the drift period equation. (The ordinate scale is -30 to +30 mK.) The set of data points on the right of the figure indicates that the final drift period begins about 150 to 180 minutes after ignition. A correction is made for the heat exchange between the calorimeter

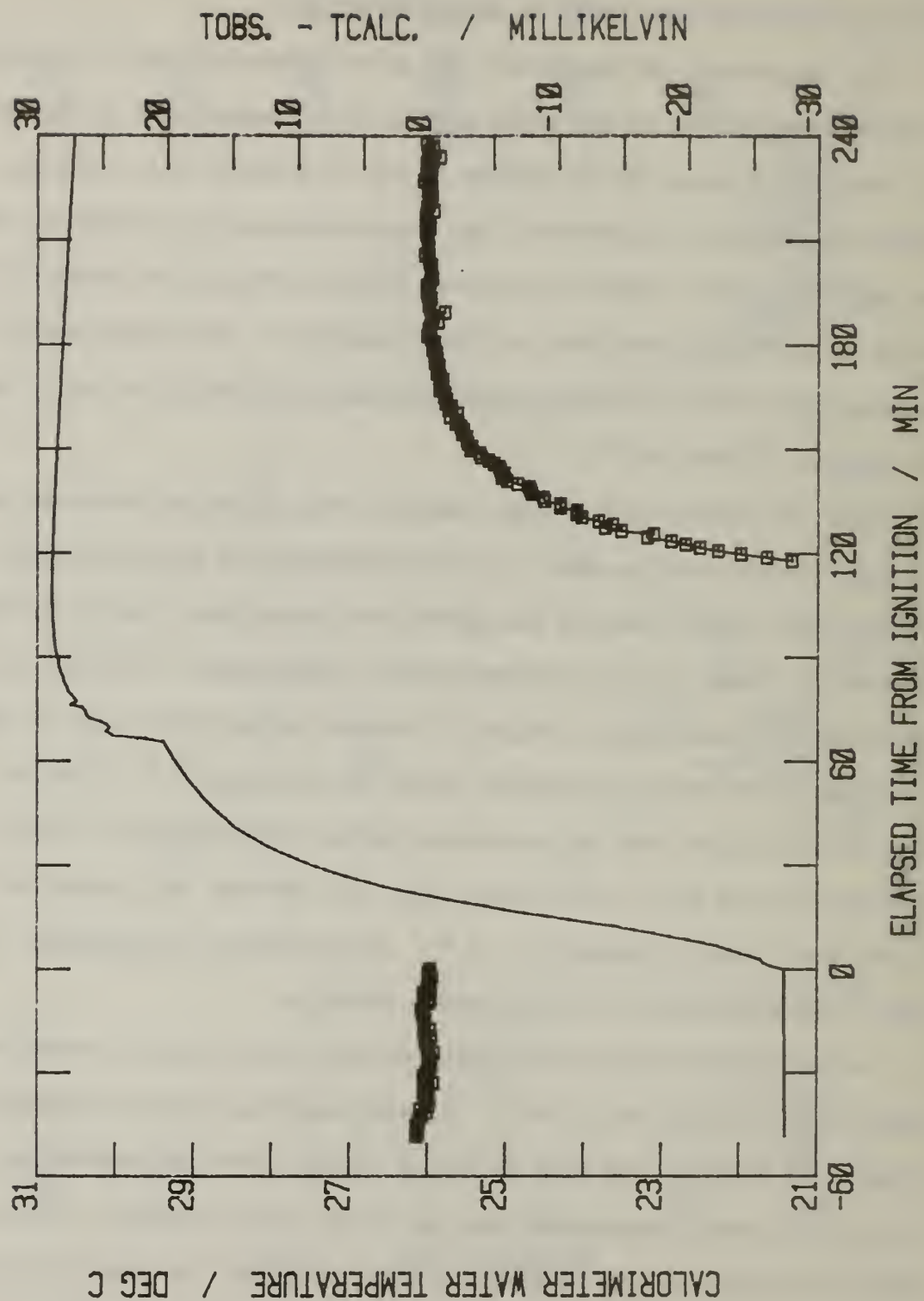


Figure 3. A plot of the temperature rise of the calorimeter water plus the deviation of the observed data points from the smooth curve fit to the data during the drift periods.

proper and the jacket during the three hour time interval between the two drift periods. This correction amounts to about 4% of the observed calorimeter water temperature rise which was between 6.0 and 9.5 °C for all experiments.

3. Sample Preparation and Measurement Protocol

To resolve the dispute about the validity of the current method of determining the calorific value of MSW, two problems were addressed. The first is the sampling problem: How does one obtain a representative gram-size sample from a kilogram-size quantity of RDF-4? The second is the processing problem: Does processing MSW down to 2 mm particle size (RDF-4) alter its calorific value significantly (i.e. one percent or more)?

Both the sampling and processing problems were examined in context of the variability of MSW for two weeks, a time period appropriate to a specification test of an incinerator-boiler system. Thirteen to twenty-three kilograms of RDF-2 were obtained on each of ten consecutive working days (5 days per week for a 2 week period from January 10 to 21, 1983) from the Baltimore County Resource and Recovery Facility in Cockeysville, Maryland. Each sample was removed from the conveyor belt immediately beyond the primary shredder before any further processing occurred. The material was brought to NBS each day, dried for 12 hours at 105 °C to determine its initial moisture content, and then stored at 5 °C.

Ideally, the processing problem could be solved by comparing the calorific value of identical pairs of kilogram-size samples of dried RDF-2, one unprocessed and the other processed. We attempted to prepare such identical kilogram-size samples by sorting each day's material into eight categories and then reconstituting the material into 2.5 kg samples having the same relative mass composition as the whole.

A summary of the results of hand sorting plus the initial moisture content

are given in Table 2. The material was hand sorted into metal, combustible, and noncombustible fractions. Metals were separated into magnetic and nonmagnetic components. The combustible fraction was separated into wood/vegetable matter, textiles, plastics, and paper. The noncombustible fraction was divided into glass/ceramics and "fines" (primarily grit and sand). The compositions according to main groups are given in lines 10 through 12 of Table 2. Metals were omitted from the reconstituted material. The combustible and noncombustible compositions of the samples actually burned in the flow and bomb calorimeters are listed in the last two rows of Table 2. The sorting and reconstituting of the material for each day required the efforts of three or more people for approximately one day.

Four of the ten day's material were selected for our study: days 6 and 9, which had the largest and smallest noncombustible content, respectively, and days 4 and 7, which had intermediate noncombustible contents. The bottom row of each of these days is underlined in Table 2.

Processed MSW consists of a low density component, fluff, which resembles chopped up cotton fibers, and a high density component which resembles sand or metal particles. Some high density material is entrained in the fluff; this entrainment is not uniform. Sampling problems arise because of segregation of the high and low density components. The sampling technique used in the first NBS study [3] of the variability of MSW for a two week period in January 1980 consisted of blending (in drums filled with steel balls), coning and quartering, and taking gram-size grab samples. This method was found to be unsatisfactory. In that study, the percent coefficient of variation⁽⁺⁾, %CV, of the enthalpy of

 (+) $100 \times (s / \text{average})$, where s is the standard deviation of a single measurement.

Table 2. Composition of municipal solid waste obtained Jan 10-21, 1983, from the Baltimore County Resource and Recovery Facility.

Day	Mass Percent (Dry)									
	1	2	3	4	5	6	7	8	9	10
1) Initial Moisture	35.0	28.9	52.7	27.6	29.3	34.3	25.1	30.8	22.0	26.4
2) Metal (Magnetic)	9.1	8.6	6.4	5.0	5.7	6.0	2.6	0.3	2.2	3.9
3) Metal (Nonmagnetic)	1.6	1.3	1.4	1.3	1.9	1.5	1.3	0.9	0.6	1.5
4) Wood, Vegetable	1.4	0.5	0.7	0.8	2.2	1.3	0.4	0.2	2.5	1.5
5) Textiles	2.8	0.6	0.5	1.6	3.2	3.2	5.2	11.4	13.8	4.9
6) Plastics	5.3	4.6	4.6	3.3	5.4	6.2	8.2	12.8	6.6	6.2
7) Paper	58.2	63.0	62.2	66.2	57.9	55.8	67.7	58.6	64.3	69.8
8) Glass, Ceramics	1.0	3.4	2.6	2.6	3.5	2.4	0.8	0.5	0.3	0.9
9) "Fines"	20.6	18.0	21.6	19.2	20.2	23.6	13.8	15.3	9.7	11.3
- - - -	-	-	-	-	-	-	-	-	-	-
10) Metals	10.7	9.9	7.8	6.3	7.6	7.5	3.9	1.2	2.8	5.4
11) Combustible Content	67.7	68.7	68.0	71.9	68.7	66.5	81.5	83.0	87.2	82.4
12) Noncombustible Content	21.6	21.4	24.2	21.8	23.7	26.0	14.6	15.8	10.0	12.2
- - - -	-	-	-	Samples as Burned			-	-	-	-
13) Combustible Content	75.8	76.2	73.8	76.7	74.3	71.9	84.8	84.0	89.7	87.1
14) Noncombustible Content	24.2	23.8	26.2	<u>23.3</u>	25.7	<u>28.1</u>	<u>15.2</u>	16.0	<u>10.3</u>	12.9

combustion on a dry basis was 3.2% for supposed identical gram-size samples of RDF-4 burned in a bomb calorimeter. The corresponding quantity for a pure (or truly homogeneous) material should be of the order of 0.1%.

For this study, we introduced a new technique for obtaining gram-size test samples. A Brinkmann Co. Sample Divider PTZ⁽⁺⁾ was used. It consists of a central cup from whose bottom surface tubes extend outward and downward to collector bottles. In the model we used there are eight tubes and eight bottles. The cup, tubes, and bottles rotate around the central vertical axis of the cup. The sample is introduced into the divider through a funnel feeding a horizontal vibrating trough which teases the sample into the divider cup.

An initial assessment of the new sample division technique showed this method to be superior to the coning and quartering procedure used in the variability study [3]. This conclusion is confirmed by the results cited in section 4.

The measurement protocol adopted for our study is summarized in Table 3. Four bags of the reconstituted kilogram-size samples of RDF-2 were selected from each of the four days. The contents of two of the bags, designated A and B, were burned in the large flow calorimeter (Step 2C, Table 3) without further processing after drying to determine the residual moisture (Step 2A). A known amount of water was added to each of the samples and then each sample was compacted to form a pellet (Step 2B). The water was added to improve the cohesion of the pellet.

(+) The commercial sources cited in this paper are included to adequately describe the experimental procedures. Such identification does not imply recommendation or endorsement by the National Bureau of Standards.

Table 3. Measurement protocol.

1. Select four 2.5 kg samples from each day.
For example:
Day 4, label samples A⁴, B⁴, C⁴, D⁴.
2. A⁴, B⁴ A) Dry.
 B) Add water and pelletize.
 C) Burn in flow calorimeter.
3. C⁴, D⁴: A) Reduce to 2 mm particle size.
 B) Blend, take grab sample (40 grams).
 C) Divide to obtain test samples (2 grams each).
 D) Recombine remainder of C⁴ and D⁴ to obtain
 two identical kilogram-size samples.
 Label E⁴, E'⁴.
 E) Test samples: two sets of
 two combustion bomb calorimeter,
 two residual moisture, and
 two furnace ash determinations
 (one set for each bag).
4. E⁴, E'⁴: A) Pelletize.
 B) Burn in flow calorimeter.

The contents of the remaining two bags, designated C and D, were processed to 6 mm sieve size using a Williams Hammer Mill and then to 2 mm sieve size using a Wiley Mill (Step 3A). A diagram of the measurement protocol for bags C and D is given in Figure 4. Powdered dry ice was mixed with the material in the throat of the Wiley Mill to prevent excessive local heating. The resulting RDF-4 of each bag was then blended in a vee blender and a forty-gram grab sample was removed (Step 3B). (The grab sample was saved for future use in another study.) Gram-size test samples were then obtained from the milled, blended material of bag C and bag D using the Brinkmann Sample Divider (Step 3C). In Step 3D, the remainder of each bag was divided into two equal parts. One part of bag C and one part of bag D were combined to form the sample of bag E and similarly for identical bag E'. Kilogram-size pellets of RDF-2 (Step 2B) and RDF-4 (Step 4A) were prepared by placing the sample in a die piece and compressing it with a total force of 534 kN.

Masses of samples were recorded before, after, and during each step of the measurement protocol to account for sample loss. Individual measurements in sets made with either the bomb or flow calorimeter were carried out in a random order.

In subsequent sections of this paper, the labels C and D refer to the gram-size test samples withdrawn from bags C and D, respectively, before step 3D in which the remainder of bags C and D were split and recombined to obtain bags E and E'. The strategy behind step 3D was to take further advantage of the improved capability of the Brinkmann Sample Divider to divide the sample into identical parts. Having prepared identical gram-size samples in Step 3C, we prepared identical kilogram-size samples in Step 3D. Hence, differences in the enthalpy of combustion of each pair of E and E' samples for a given day afford a direct measure of the measurement precision of the new calorimeter on an actual

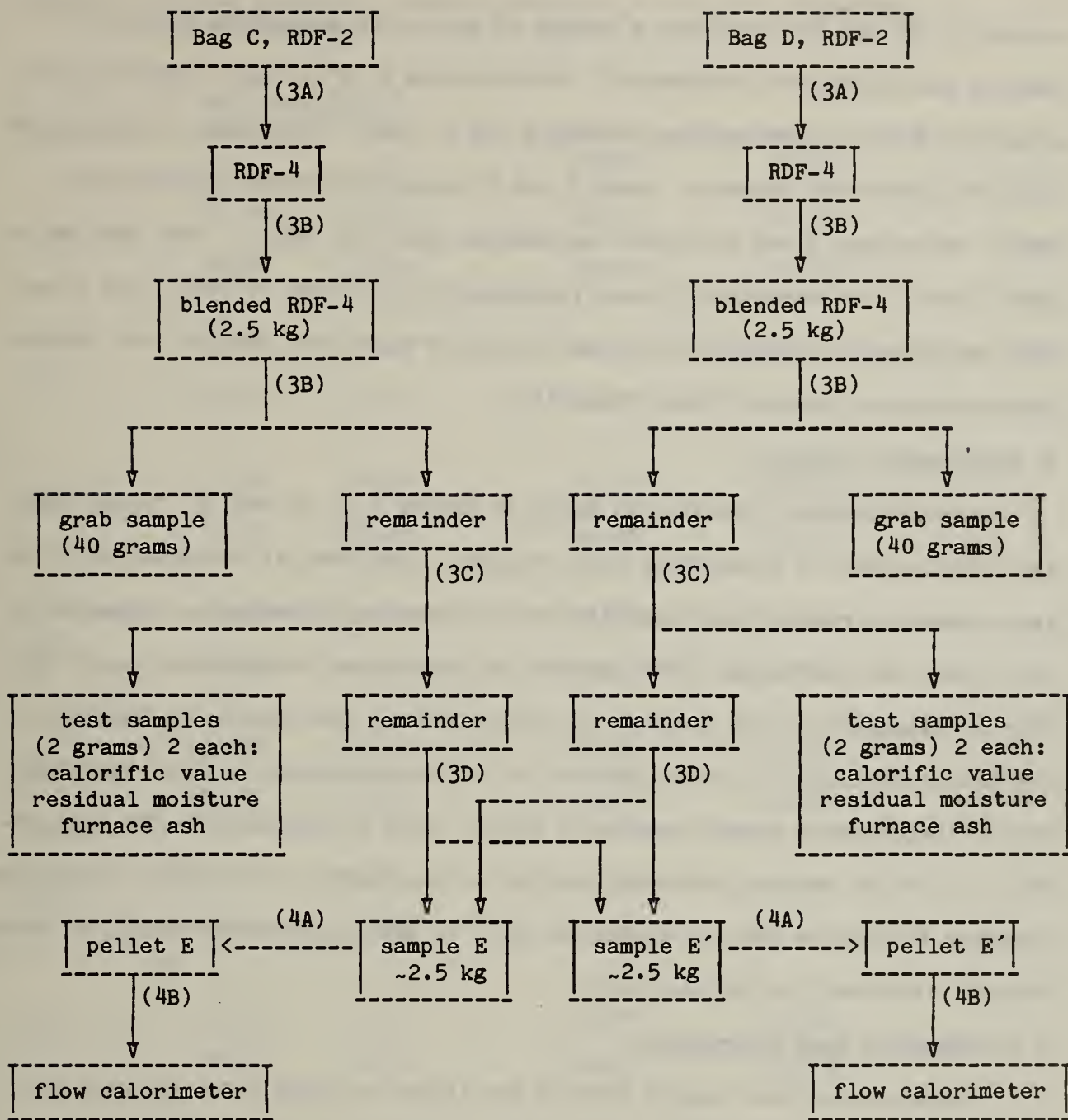


Figure 4. Diagram of the sample division technique used to prepare gram-size test samples (C and D) for the bomb calorimeter studies and kilogram-size samples (E and E') for the flow calorimeter studies from 2.5 kilogram hand sorted and reconstituted samples in bags C and D. The steps given in Table 3 are indicated in the parentheses along the arrows.

sample of RDF (as distinct from a sample of pure cellulose which does not contain noncombustible components). While labels E, E', A and B refer to kilogram-size samples, the samples of bags E and E' differ from those of bags A and B in two important respects. Bags E and E' contain processed material (i.e. RDF-4) while bags A and B contain unprocessed MSW (i.e. RDF-2). The samples in bags E and E' are essentially exact replicates while those in bags A and B are only approximately identical (i.e. to the extent that hand sorting into general classifications given in Table 2 permits).

4. Experimental Results

The experimental results are given in Tables 4, 6, 7, and 9. Three important conclusions are drawn from these results. The first is concerned with the improvement in results obtained with the new sampling technique as compared to the coning and quartering technique used in the earlier variability study [3]. This is discussed in the analysis of the gram-size test results in Section 4.1. The second conclusion concerns whether or not gram-size test samples represent their kilogram-size parent samples of RDF-4. This is discussed in the comparison of flow and bomb calorimetry results in Section 4.2.1. The final conclusion concerns whether or not processing of RDF-2 to RDF-4 alters the calorific value. This is discussed in Section 4.2.2.

4.1 Combustion Bomb Calorimetry

The gram-size test sample results are listed in Table 4 for sampling days 4, 6, 7, and 9. They consist of duplicate results for each sample of the enthalpy of combustion (ASTM Method E711-81), bomb ash (dried residue from a combustion bomb experiment) and furnace ash (ASTM Method E830-81). All results were converted to a dry basis using moisture contents obtained with ASTM Method E790-81. In the furnace ash test, the sample is combusted in air in a furnace at 575 °C for two hours.

Table 4. Ash and combustion bomb calorimetry results for two-gram test samples.

Sample	Ash (Dry Basis)		HHV2 MJ/kg	HHV3-B MJ/kg
	Furnace %	Bomb %		
C4	34.45	33.62	14.619	22.022
	34.54	33.50	14.555	21.886
D4	32.73	31.54	14.867	21.716
	32.30	31.21	14.857	21.598
C6	37.82	34.58	14.678	22.436
	37.19	35.67	14.546	22.612
D6	37.67	36.18	14.509	22.735
	36.89	35.51	14.714	22.817
C7	27.76	26.78	16.273	22.224
	28.30	26.58	16.506	22.480
D7	25.04	24.19	17.254	22.760
	25.49	23.13	17.253	22.443
C9	17.89	16.41	18.200	21.772
	17.68	16.31	18.221	21.773
D9	17.43	15.84	18.336	21.787
	17.55	16.57	18.325	21.966

Table 5 summarizes the statistical analysis of the results; the methods are summarized in references [5] and [6]. The percent coefficient of variation, %CV, for the between-day, within-day, and within-bag variability are calculated for furnace ash, bomb ash and calorific values. HHV2 and HHV3 are the negative of the enthalpy of combustion on a moisture free and a moisture-ash free basis, respectively. HHV3-B refers to the use of the bomb ash value in the calculations; HHV3-F refers to the use of the furnace ash value in the calculations. For sake of comparison, the corresponding results for the 1980 variability study are given in the lower half of Table 5. The total %CV in column 6 is the square root of the sum in quadrature of the %CV's given in the preceding three columns.

In this work, the within-bag %CV, which corresponds to the combined measurement and sampling error for gram-size samples, is substantially less for all properties than in the prior variability study [3]. The small value for the HHV2 within-bag %CV and its close agreement with the within-bag values for HHV3-B and HHV3-F indicates that the new sample division procedure has reduced our sampling error for gram-size samples to less than 0.8%. The within-day %CV for HHV2, 2.0%, suggests that the extra error introduced in failing to prepare equivalent kilogram-size samples (C and D) for each day by hand sorting and recombination is somewhat larger in size than the combined measurement and sampling error. (The label "within-day %CV" is the between-bags %CV estimated within days.)

The fact that the within-day and within-bag %CV for HHV3-B are about the same while the corresponding values for HHV2 are not indicates that the hand sorting and reconstituting procedure failed to yield equivalent samples for each day primarily because the initial noncombustible contents were different. The lower between-day %CV for HHV3-B (or HHV3-F) as contrasted to that for HHV2 suggests that the primary variability in the MSW stream is in its ash (and

Table 5. Summary of statistical analysis of ash and combustion bomb calorimetry results for two-gram test samples.

Property (Dry basis)	Components of Variability, %CV				
	Average	Within Bag	Within Day	Between Days	Total
January 1983 Study (4 days, this work)					
Furnace Ash, %	28.8	1.2	4.1	29.9	30.2
Bomb Ash, %	27.4	1.7	4.8	31.1	31.5
HHV2, MJ/kg	16.108	0.5	2.0	10.9	11.1
HHV3-B, MJ/kg	22.189	0.6	0.6	1.9	2.1
HHV3-F, MJ/kg	22.649	0.8	0.3	2.5	2.6
January 1980 Study (10 days, [3])					
Furnace Ash, %	31.8	6.5	23.0	29.8	38.2
Bomb Ash, %	28.9	10.6	15.6	33.7	38.6
HHV2, MJ/kg	15.393	3.2	6.6	11.2	13.4
HHV3-B, MJ/kg	21.664	1.6	1.5	3.7	4.3
HHV3-F, MJ/kg	22.523	5.1	5.4	3.1	8.0

initial moisture) content and not the calorific value of its combustible content. The latter was also a conclusion of the first variability study [3].

The larger average value for HHV3-F as compared to HHV3-B is a reflection of the fact that the percentage of furnace ash is greater than the percentage of bomb ash. The data of Table 4 can be used to calculate a value of furnace ash minus bomb ash of $1.44 \pm 0.33\%$. The uncertainty, .33%, is the imprecision of the mean at the 95% confidence level⁽⁺⁾.

It should be noted that the total %CV reflects primarily the day-to-day variability and not the variability of the measurement processes being studied here. The latter is characterized by the within-bag and, to a certain extent, the within-day %CV. The fact that the between-day %CV is much larger than the within-bag or within-day %CV's in Table 5 does not mean that the between-day %CV is the limit to which a two-week supply of MSW (i.e. an incinerator test sample) can be characterized. Rather it means that each day's properties must be characterized separately and/or samples should be composited as described by Mandel and Paule [5].

It should also be noted that neither the within-day or between-day %CV of this work are strictly comparable to the corresponding parameters of the earlier variability study [3]. The values of the within-day %CV of this work have been deliberately altered by the hand sort and reconstitution procedure. The between-day %CV is for only four of ten days samples of MSW and the four days were not selected at random.

(+) For this and all other statements labeled "imprecision", we are using the product of the standard deviation of the mean and the appropriate Student t factor at the 95% confidence level.

4.2 Combustion Flow Calorimetry

Six determinations of the energy equivalent (i.e. apparent heat capacity) of the flow calorimeter gave a %CV for a single measurement of 0.26%. The average energy equivalent was 4.509 MJ/K with an imprecision of ± 0.012 MJ which is 0.27% ($0.26\% \times \text{Student } t / \sqrt{n}$) of the mean.

In the calculation of the flow calorimetry results, we have omitted several minor calorimetric corrections (e.g. for small amounts of CO in the product gas, for small amounts of unburned sample in the ash, etc.). We do not expect the absence of these corrections to affect our results or conclusions in a practically significant manner. For convenience the results of the flow calorimetry will, henceforth, be referred to as flow results.

4.2.1 Flow Calorimetry of RDF-4.

The flow and bomb results on processed MSW (RDF-4) are summarized in Table 6. The columns labeled furnace ash, bomb ash, bomb HHV2, bomb HHV3 are calculated from the averages of the gram-size test measurements on C and D weighted by the amounts of sample from bags C and D which form the samples in bags E and E'. (See Table 7.)

The analysis of the data in Table 6 is given in Table 8. The first four rows show the average and imprecision of the paired differences for values of flow ash minus furnace ash and flow minus bomb values for ash, HHV2 and HHV3. The day average rather than each of the paired differences of flow minus bomb values for ash, HHV2, and HHV3 were used to evaluate imprecisions as the differences tend to group themselves into within-day sets.

The first two rows of Table 8 show that the average flow ash is $2.7 \pm 1.5\%$ greater than the average bomb ash and is $2.5 \pm 1.9\%$ smaller than the average furnace ash. These differences are consistent with the relation between the combustion zone temperatures (i.e. bomb > flow > furnace). They are also

Table 6. Flow and bomb calorimetry results on processed MSW (RDF-4).

Sample	Ash (Dry Basis), %			HHV2, MJ/kg		HHV3, MJ/kg	
	Furnace ^(a)	Flow	Bomb ^(a)	Flow	Bomb ^(a)	Flow	Bomb ^(a, b)
E4	33.51	32.78	32.48	14.732	14.723	21.918	21.804
E'4	33.50	32.94	32.47	14.838	14.725	22.126	21.804
E6	37.43	36.17	35.49	14.458	14.612	22.650	22.650
E'6	37.43	36.32	35.49	14.489	14.612	22.752	22.650
E7	26.63	26.23	25.17	16.725	16.821	22.671	22.479
E'7	26.62	26.11	25.17	16.735	16.821	22.648	22.479
E9	17.63	17.20	16.28	18.425	18.272	22.253	21.826
E'9	17.64	17.11	16.28	18.295	18.271	22.070	21.825

(a) Calculated from averages of bomb calorimetry measurements on bags C and D and known mass fractions of C and D used to make up bags E and E'.

(b) HHV3-B

Table 7. Mass Fractions of Bags E and E' contributed by Bags C and D.

Components (Dry Basis)				
Bag	Bag	Mass Fraction	Bag	Mass Fraction
E4	C4	0.50333 ^(a)	D4	0.49667
E'4	C4	0.49891	D4	0.50109
E6	C6	0.49901	D6	0.50099
E'6	C6	0.49862	D6	0.50138
E7	C7	0.50078	D7	0.49922
E'7	C7	0.50036	D7	0.49964
E9	C9	0.49043	D9	0.50957
E'9	C9	0.49478	D9	0.50522

(a) Fifth significant figure is carried for computational purposes.

Table 8. Analysis of flow and bomb calorimetry results on processed MSW (RDF-4).

Property Difference (Dry Basis)	Average	Imprecision ^(a)
1) Flow Ash-Bomb Ash, %	+0.76 (+2.7%) ^(b)	0.41 (1.5%)
2) Flow Ash-Furnace Ash, %	-0.69 (-2.5%)	0.54 (1.9%)
3) HHV2 Flow-HHV2 Bomb, MJ/kg	-0.020 (-0.1%)	0.178 (1.1%)
4) HHV3 Flow-HHV3 ^(c) Bomb, MJ/kg	+0.197 (+0.9%)	0.187 (0.8%)

Components of Variability, %CV

Property (Dry basis)	Average	Measurement Error	Between Day
5) Ash, Flow, %	28.11	0.3	30.0
6) HHV2, Flow, MJ/kg	16.087	0.4	11.3
7) HHV3, Flow, MJ/kg	22.386	0.5	1.5
8) Ash, Bomb, %	27.35		
9) Ash, Furnace, %	28.80		
10) HHV2, Bomb, MJ/kg	16.107		
11) HHV3, Bomb, MJ/kg	22.189		

(a) product of Student t for the 95% confidence level and the standard deviation of the average.

(b) percent of flow value

(c) HHV3-B

consistent with the observed variation in the amount of furnace ash formed as a function of furnace temperature [7]. The average difference of the flow minus bomb values for HHV2 is -0.1% of the average HHV2 flow value. This difference is not statistically different from zero as the 95% confidence interval is from -1.2 to +1.0%. The average difference of the flow minus bomb values for HHV3 is +0.9% of the average HHV3 flow value. The difference is statistically different from zero as the 95% confidence interval is from +0.1 to +1.7%. However, the interval misses zero by such a small amount that we do not regard the non-overlap to be practically important.

The average difference of the flow minus bomb values for HHV3 can be different from zero even if the HHV2 difference were zero, because, in this work the average difference of the flow minus bomb ash values is +0.8% which is 2.7% of the mean ash value. This causes the HHV3 difference to be +1.1% even if the HHV2 difference were zero. This is approximately what is observed in rows 3 and 4 of Table 8 (e.g. the average difference of the flow minus bomb values for HHV3 is 1.0% larger than the average difference of the flow minus bomb values for HHV2).

The difference of the flow minus bomb values for HHV2 completes the evidence for the conclusion, drawn in the previous section, that the new sample division technique provides representative gram-size test samples. The results of this section show that HHV2 is the same for the kilogram-size parent and the gram-size test samples; the results of the previous section showed that gram-size test samples from the same bag were practically the same.

The measurement errors given in the fifth through seventh rows of Table 8 are obtained from the analysis of the flow ash, HHV2, and HHV3 using paired differences from replicate samples E and E' for each day. Between-day variability was computed in the usual manner [6]. The precision of a calorimetric

measurement on kilogram-size samples of dry, processed MSW (0.4 to 0.5%) and on dry cellulose (0.3%) are comparable in magnitude. Confirmation of the assertion that E and E' are replicate samples is obtained from the fact that the measurement error from the flow ash is small.

4.2.2 Flow Calorimetry of RDF-2

The flow results on unprocessed MSW (RDF-2) are summarized and compared with those on processed MSW (RDF-4) in Table 9. Results on the unprocessed samples A and then B for each day are given in columns headed "UNPR"; those on the processed samples E and then E' for each day are given in columns headed "PR". An analysis of the data in Table 9 is given in Table 10. Average values and imprecisions in the differences of the unprocessed minus processed values for ash, HHV2, and HHV3 are given in the first three rows of Table 10. All differences have 95% confidence intervals that overlap zero and, thus, are not significantly different from zero (at this confidence level). The average effect of processing kilogram-size samples of RDF-2 to RDF-4 is such as to increase HHV2 by 0.5% and leave HHV3 unchanged. The 95% uncertainty intervals for the HHV2 difference, -3.4 to +2.4%, and the HHV3 difference, -2.0 to +2.0%, are larger than those for the differences of the flow minus bomb values given in Table 8 because hand sorting and reconstituting, as we carried it out, did not produce completely identical bags.

Our hand sorting process was subjective because personnel assisting in this task varied and composite materials (e.g. insulated wire or some types of shoes) were not separated into components prior to being assigned a category. Moreover, finer distinctions that might be significant calorimetrically were not made (e.g. distinction between hard and soft plastics or cotton and artificial fiber textiles). It is interesting to note that the day-averages of the percent ash content listed in Table 6 bear the same relation to each other as the total

Table 9. Flow calorimetry results on unprocessed (UNPR) and processed (PR) MSW.

Day	Ash (Dry Basis), %		HHV2, MJ/kg		HHV3, MJ/kg	
	UNPR ^(a)	PR ^(b)	UNPR ^(a)	PR ^(b)	UNPR ^(a)	PR ^(b)
4	33.19	32.78	14.370	14.732	21.507	21.918
	34.39	32.94	14.306	14.838	21.806	22.126
6	36.06	36.17	14.432	14.458	22.570	22.650
	39.00	36.32	14.315	14.489	23.466	22.752
7	26.07	26.23	16.976	16.725	22.961	22.671
	24.00	26.11	17.039	16.735	22.418	22.648
9	16.97	17.20	18.350	18.425	22.100	22.253
	18.06	17.11	18.250	18.295	22.272	22.070

(a) results for each day are sample A and then B.

(b) results for each day are sample E and then E'.

Table 10. Analysis of flow calorimetry results on kilogram-size samples of processed and unprocessed MSW.

Unprocessed-Processed Property Differences (Dry Basis)	Average	Imprecision ^(a)
1) Ash, %	+0.36 (+1.3%) ^(b)	1.70 (6.0%)
2) HHV2, MJ/kg	-0.083 (-0.5%)	0.471 (2.9%)
3) HHV3, MJ/kg	+0.001 (+0.0%)	0.446 (2.0%)

Components of Variability, %CV			
Unprocessed Property (Dry Basis)	Average	Within Day	Between Day
4) Ash, %	28.47	4.9	31.4
5) HHV2, MJ/kg	16.005	0.3.	12.3
6) HHV3, MJ/kg	22.388	1.7	2.3

(a) product of Student t for the 95% confidence level at 3 degrees of freedom and the standard deviation of the average

(b) percent of average

percent of nominally noncombustible contents listed in Table 2, line 14. However, the percent ash content is larger in absolute magnitude than the value given in line 14 by $9.1 \pm 2.4\%$. This is consistent with the fact that all components of our nominally combustible category produce some ash.

Within-day and between-day components of variability of the data for unprocessed samples are given in rows four to six of Table 10. Comparing these values with the corresponding bomb values listed in Table 5 shows the within-day %CV for HHV2 and HHV3 are in general agreement.

4.3 Summary of Statistical Results

(1) Replicate kilogram-size samples (i.e. samples E and E') of processed MSW (RDF-4) have been prepared for the first time by a new sample division technique.

(2) The differences in the enthalpy of combustion (on a dry basis) of replicate samples E and E' provided a %CV for a single measurement with the new flow calorimeter on an actual sample of RDF of 0.4%. (See Table 8.)

(3) The new sample division technique provided gram-size test samples, from the same kilogram-size parent sample, that have a %CV for a single measurement of 0.5% for the enthalpy of combustion (on a dry basis). (See Table 5.)

(4) The new sample division technique provided gram-size test samples from kilogram-size samples of RDF-4 which have the same average enthalpy of combustion (on a dry basis) as the kilogram-size parent sample to within 0.1%. The 95% confidence interval for the average difference in the enthalpy of combustion of parent minus gram-size test samples ranges from -1.2 to +1.0%. (See Table 8, line 3.)

(5) Processing of kilogram-size samples of MSW having a sieve size of minus 15 cm (RDF-2) down to a 2 mm sieve size (RDF-4) alters the enthalpy of combustion of the material (on a dry basis) by less than 0.5%. The 95% confidence

interval for the average difference of the unprocessed minus processed values is from -3.4 to 2.4%. The main contribution to this confidence interval is the variability in the supposed identical samples prepared by hand sorting. (See Table 10.)

(6) Kilogram-size samples of dried RDF-2 having the same enthalpy of combustion (on a dry basis) to about 2% can be prepared by matching the relative mass compositions according to the following categories of: metals, textiles, plastics, wood/vegetable matter, paper, glass/ceramics, and "fines" (residual grit and sand). (See Tables 5 and 10.)

5. Discussion and Conclusions

As indicated in the introduction, this work was performed as part of the effort to resolve the general question of whether or not the calorific value of a multiton quantity of MSW can be obtained by sampling down to one or two grams and obtain representative test samples. This general question can be divided into two parts. The first is: How should one sample a multiton pile or stream of raw waste to obtain a number of kilogram-size increments of sufficient number and amount to characterize the calorific value with a given precision? The second is: How can one obtain representative gram-size test samples that characterize these increments?

The first question is currently being investigated by ASTM Committee E-38⁽⁺⁾ and is beyond the scope of this paper.

The sample preparation and measurement protocol of this work were designed

(+) The subject and principles of the design of a sampling protocol, based on experimental test, are addressed in ASTM Method D 2234-76, Standard Method for the Collection of a Gross Sample of Coal. A variant of this method, as applied to MSW, is under study by the ASTM Committee E-38.

primarily to answer the second question. We assumed that the sample increment was, or could be, conveniently divided into bags containing between 2 and 2.5 kg of minimally processed MSW (RDF-2). To answer the second question, two problems were studied.

The first problem was the processing problem: Does processing of RDF-2 (MSW of 150 mm or less particle size) down to RDF-4 (MSW of 2 mm or less particle size) alter the calorific value of the waste?

The second problem was the kilogram-to-gram sampling problem: How does one obtain a representative gram-size test sample from a kilogram quantity of RDF-4?

Our results show that processing from 150 mm to 2 mm particle size does not alter the calorific value of the MSW within the uncertainty of our experiments. The average of the unprocessed minus processed difference in the calorific value on a dry basis is -0.5% and on a dry, ash-free basis is 0.0% of the corresponding mean calorific value. The imprecision in these differences are 2.9 and 2.0%, respectively of the corresponding mean calorific value. Most of the imprecision is due to the fact that completely identical pairs of samples were not prepared by our hand sorting and reconstituting of RDF-2. An important qualification is that excessive local heating of the RDF-2 during milling was avoided. In our case, powdered dry ice was added to the sample as is it was fed to a Wiley Mill in the final step of processing.

In regards to the kilogram-to-gram sampling problem, our results show that it is possible to obtain gram-size test samples that have the same calorific value as the kilogram parent of RDF-4 if one uses the new sample division technique described in this work. We show that different gram-size test samples taken from a single parent kilogram-size RDF-4 sample have the same calorific value as the parent within the uncertainty of our experiments. The average difference of the parent minus gram-size sample calorific values is -0.1% on a

dry basis and +0.9% on a dry, ash-free basis of the corresponding mean calorific value. The imprecisions of the differences are 1.1 and 0.8%, respectively, of the corresponding mean calorific values.

The new sample division technique was successful because it is insensitive to the effects of segregation whereas our earlier coning and quartering method was not. Because the technique employed in this study is time consuming, it is recommended only as a reference procedure. An apparatus with more rapid throughput is needed for routine testing. We definitely do not recommend coning and quartering plus taking of grab samples as a technique for the preparation of gram-size test samples of RDF-4. However, it can be used if a within-bag %CV in the calorific value (on a dry basis) of three percent or more is tolerable.

In summary, this work shows that if 2.5 kg increments are collected to characterize the calorific value of a MSW stream, each increment can be processed (provided that excessive local heating is avoided) and then be sampled using the sample division technique cited here (or any other procedure sufficiently insensitive to segregation) to obtain representative gram-size test samples. Combining⁽⁺⁾ the results on the processing to small particle size (see Table 10) and kilogram-to-gram sampling problem (see Table 8), we show that the average difference in the calorific values of the minimally processed (RDF-2)

(+) Averages of the results on each basis (i.e. dry or dry, ash-free) were obtained by adding; the imprecisions were obtained as the square root of the sum of the squares, i.e.

$$\text{Dry: } (-0.5) + (-0.1) \pm \sqrt{(2.9)^2 + (1.1)^2} \%$$

$$\text{Dry, ash free: } (0.0) + (+0.9) \pm \sqrt{(2.0)^2 + (0.8)^2} \%$$

kilogram-size sample minus the processed (RDF-4) gram-size sample is -0.6% on a dry basis and +0.9% on a dry, ash-free basis of the corresponding mean calorific values. The imprecisions in these values are 3.1 and 2.2%, respectively, of the mean calorific values.

6. References

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11. ABSTRACT (A 200-word or less factual summary of most significant information. If document includes a significant bibliography or literature survey, mention it here) <p>A study has been made at the National Bureau of Standards to establish the limits of reliability of the calorific value of municipal solid waste (MSW) by the bomb calorimetric procedure currently used in commercial test laboratories. This procedure involves using gram-size samples derived from MSW that has been processed down to a particle size of 2 mm or less. Critics of the procedure argue that gram-size samples are too small to be representative of such a heterogeneous material, and that processing MSW alters its composition.</p> <p>To test the bomb calorimetric procedure, a 2.5 kg capacity combustion flow calorimeter was designed and constructed for the determination of the enthalpies of combustion of kilogram-size samples of MSW in flowing oxygen near atmospheric pressure.</p> <p>Calorimetric data on processed MSW were obtained using both the large flow and a gram-size bomb calorimeter. Intercomparison of results shows that the calorific value (on a dry basis) of gram-size test samples agrees, within the uncertainty of our experiments, with the corresponding values for their kilogram-size parent samples provided that the sample division technique used to obtain the gram-size samples is that described in this work. The average difference of the parent minus gram-size sample values (on a dry basis) is -0.1% with an imprecision (95% confidence level) of $\pm 1.1\%$ of the mean calorific value.</p> <p>The effects of processing on sample composition were determined by (continued)</p>			
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intercomparison of flow calorimetric results on kilogram-size samples of processed and minimally processed MSW (150 mm or less particle size) that are nominally identical. The average difference of the unprocessed minus processed values (on a dry basis) is -0.5% with an imprecision (95% confidence level) of $\pm 2.9\%$ of the mean calorific value.

